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## Physicochemical and Spectroscopic Characterization of *p*-Chlorobenzaldehyde: An Impact of Biofield Energy Treatment

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### Abstract

*p*-Chlorobenzaldehyde (*p*-CBA) is used as an important chemical intermediate for the preparation of pharmaceuticals, agricultural chemicals, dyestuffs, optical brighteners, and metal finishing products. The study aimed to evaluate the effect of biofield energy treatment on the physicochemical and spectroscopic properties of *p*-CBA. The study was accomplished in two groups *i.e.* control and treated. The control group was remained as untreated, while the treated group was subjected to Mr. Trivedi's biofield energy treatment. Finally, both the samples (control and treated) were evaluated using various analytical techniques. The surface area analysis showed a substantial increase in the surface area by 23.06% after biofield treatment with respect to the control sample. The XRD analysis showed the crystalline nature of both control and treated samples. The X-ray diffractogram showed the significant alteration in the peak intensity in treated sample as compared to the control. The XRD analysis showed the slight increase (2.31%) in the crystallite size of treated sample as compared to the control. The TGA analysis exhibited the decrease (10%) in onset temperature of thermal degradation from 140°C (control) to 126°C in treated sample. The  $T_{max}$  (maximum thermal degradation temperature) was slightly decreased (2.14%) from 157.09°C (control) to 153.73°C in treated sample of *p*-CBA. This decrease in  $T_{max}$  was possibly due to early phase of vaporization in treated sample as compared to the control. The FT-IR spectrum of treated *p*-CBA showed the increase in wavenumber of C=C stretching as compared to the control. The UV spectroscopic study showed the similar pattern of wavelength in control and treated samples.

Altogether, the surface area, XRD, TGA-DTG and FT-IR analysis suggest that Mr. Trivedi's biofield energy treatment has the impact to alter the physicochemical properties of *p*-CBA. This treated *p*-CBA could be utilized as a better chemical intermediate than the control *p*-CBA for the synthesis of pharmaceutical drugs and organic chemicals.

**Keywords:** Biofield energy treatment; *p*-Chlorobenzaldehyde; Surface area; X-Ray diffraction; Fourier transform infrared spectroscopy; UV-Vis spectroscopy.

**Abbreviations:** NCCAM: National Center for Complementary and Alternative Medicine; NIH: National Institute of Health; XRD: X-ray diffraction; TGA: Thermogravimetric analysis; DTG: Derivative Thermogravimetry

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### Introduction

*p*-Chlorobenzaldehyde (*p*-CBA) is an organic compound comprising of benzene ring with formyl and chlorine substituents

at 1 and 4 positions, respectively. The *p*-CBA is used as an important reaction intermediate for the manufacturing of several pharmaceutical drugs and agricultural chemicals [1]. It is used in the production of triphenyl methane and related dyes. It is also

used for optical brighteners and metal finishing products [1,2]. The *p*-CBA is commercially produced by side-chain chlorination of *p*-CBA followed by acid hydrolysis [3]. The *p*-CBA along with *p*-chloroaniline is used for the synthesis of Schiff base. The Schiff bases are versatile imine (C=N) containing compounds having broad spectrum of biological activities [4]. The incorporation of metals in Schiff base in the form of metal complexes exhibited some degree of biological activities like antifungal [5], antibacterial [6], anticancer [7], and anti-inflammatory activity [8]. As *p*-CBA used as an important reaction intermediate, its rate of reaction plays a crucial role. The literature suggests that any alteration in crystallite size and surface area can affect the kinetics of reaction [9]. Moreover, the thermal properties *i.e.* vaporization temperature, decomposition temperature of chemical compound also affect the reaction kinetics [10]. Therefore, considering the importance of *p*-CBA, it is important to discover an alternate and safe approach, which can improve the overall physicochemical properties of compound. Recently, biofield energy treatment has been reported to alter the physicochemical as well as spectral properties of various organic compounds and pharmaceutical drugs [11-13].

The National Institute of Health/National Center for Complementary and Alternative Medicine (NIH/NCCAM) conceived the healing energy (putative energy fields) treatment under the subcategory of energy therapies [14]. It is reported that human body is permeated and surrounded with the bioenergetic field (subtle energy field) [15]. The health of living organism depends on the balance of this bioenergetics field. In the diseased situation, this bioenergetics field gets depleted [16]. The experts of energy medicine manipulate and balance this bioenergetics field *via* harnessing the energy from the Universe [17]. Thus, the human (expert of energy medicine) has the ability to harness the energy from the Universe and transfer it to any living or nonliving object to balance or re-pattern the electromagnetic energy field [18]. The objects always receive this energy and respond in the useful way [19]. The biofield energy therapy is being practiced in the form of healing therapy or therapeutic touch throughout the world and especially in the western countries [20,21]. It is estimated that about 36% of Americans regularly uses some form of Complementary and Alternative Medicine (CAM) [22]. The biofield energy treatment is effectively used to stimulate the overall health of human being by reducing the pain and anxiety [23,24].

Mr. Trivedi is well known for his unique biofield energy treatment (The Trivedi Effect<sup>®</sup>) that has been evaluated in numerous arenas like agricultural research [25], biotechnology research [26], microbiology research [27,28], pharmaceutical sciences [13,19], and materials science [29,30].

Hence, based on the prominent impact of biofield energy treatment and significance of *p*-CBA as a chemical intermediate, the present study was aimed to evaluate the effect of Mr. Trivedi's biofield energy treatment on the physicochemical and spectroscopic properties of *p*-CBA. The analysis was done using surface area analyzer, X-ray diffractometry (XRD), thermogravimetric analysis-derivative thermogravimetry (TGA-DTG), Fourier transform infrared (FT-IR) spectroscopy, and UV-Vis spectrometry.

## Materials and Methods

### Study design

The *p*-chlorobenzaldehyde (*p*-CBA) was purchased from Loba Chemie Pvt. Ltd., India. The *p*-CBA was divided into two groups *i.e.* control and treated. The control sample was kept without treatment, while the treated sample in sealed pack was handed over to Mr. Trivedi to render the biofield energy treatment under laboratory conditions. Mr. Trivedi provided the biofield energy treatment to the treated group *via* his unique energy transmission process without touching the sample [13]. Afterward, both the control and treated samples were analyzed with respect to physicochemical and spectroscopic properties using various techniques like surface area analyzer, XRD, TGA-DTG, FT-IR and UV-vis spectroscopy.

### Surface area analysis

The surface area of control and treated *p*-CBA was analyzed using the Brunauer–Emmett–Teller (BET) surface area analyzer (Smart SORB 90) based on the ASTM D 5604 method. The range of the instrument was 0.2 m<sup>2</sup>/g to 1000 m<sup>2</sup>/g. The percent change in surface area was calculated with the help of following equation:

$$\% \text{ change in surface area} = \frac{[S_{\text{Treated}} - S_{\text{Control}}]}{S_{\text{Control}}} \times 100$$

Here,  $S_{\text{Control}}$  is the surface area of the control sample and  $S_{\text{Treated}}$  is the surface area of treated sample.

### XRD study

The XRD analysis of *p*-CBA (control and treated) samples was done on Phillips (Holland PW 1710) X-ray diffractometer with copper anode and nickel filter. The wavelength of XRD system was set to 1.54056 Å. The percent change in average crystallite size ( $G$ ) was calculated using following equation:

$$G = [(G_t - G_c) / G_c] \times 100$$

Here,  $G_c$  and  $G_t$  are average crystallite size of control and treated powder samples, respectively.

### TGA-DTG analysis

The TGA-DTG analysis was carried out on Mettler Toledo simultaneous TGA-DTG analyzer. The analytes were heated up to 400°C from room temperature at the heating rate of 5°C/min under air atmosphere. The onset temperature of thermal degradation and  $T_{\text{max}}$  (temperature at which maximum weight loss occur) in samples were obtained from TGA-DTG thermogram.

### Spectroscopic studies

The treated sample of *p*-CBA was divided into two groups *i.e.* T1 and T2 for the FT-IR and UV-vis spectroscopy. The spectral data of treated samples were compared with the respective spectral data of control sample.

### FT-IR spectroscopic characterization

The FT-IR spectroscopy was done to determine the effect of biofield energy treatment on dipole moment, force constant, and bond strength in chemical structure [31]. The samples for FT-IR

analysis were prepared by crushing with spectroscopic grade KBr into fine powder. Subsequently, the mixture was pressed into pellets. The spectra were obtained from Shimadzu's Fourier transform infrared spectrometer (Japan) with the frequency range of 500-4000  $\text{cm}^{-1}$ .

### UV-Vis spectroscopic analysis

The UV spectra of control and treated samples of *p*-CBA were obtained from Shimadzu UV spectrometer (2400 PC) with quartz cell of 1 cm and a slit width of 2.0 nm. The analysis was done at the wavelength range of 200-400 nm.

## Results and Discussion

### Surface area analysis

The surface area of control and treated *p*-CBA are shown in **Figure 1**. The surface area of control and treated sample were found as 0.2498  $\text{m}^2/\text{g}$  and 0.3074  $\text{m}^2/\text{g}$ , respectively. The result showed an increase in surface area by 23.06% in the treated sample with respect to the control sample. It is well known that surface area is inversely proportional to the particle size [32]. Based on this, it is presumed that biofield energy induced the milling process, which leads to decrease the particle size of treated sample. As a result, the surface area of treated sample was increased significantly.

### XRD analysis

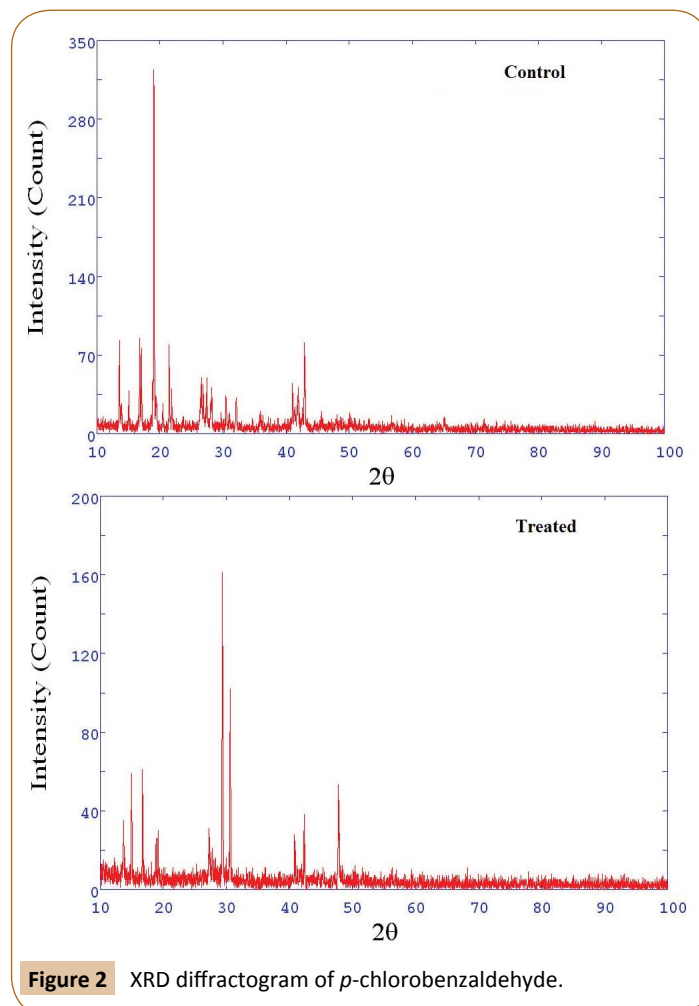
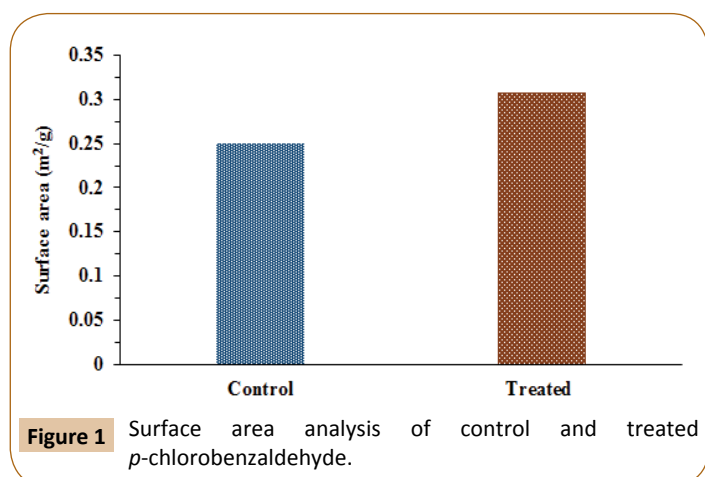
The XRD diffractograms of *p*-CBA (control and treated) samples are shown in **Figure 2**. The XRD diffractograms of both samples showed the sharp and intense peaks that suggest the crystalline nature of control and treated samples. The XRD diffractogram of control sample showed the peaks at  $2\theta$  equal to 13.6°, 16.8°, 17.09°, 18.85°, 19.06°, 21.48°, 26.86°, 27.45°, 41.01°, and 42.91°. Similarly, the XRD diffractogram of treated *p*-CBA exhibited the XRD peaks at  $2\theta$  equal to 13.73°, 14.98°, 16.71°, 19.19°, 27.26°, 29.38°, 30.61°, 40.81°, 42.35°, and 47.77°. The **Figure 2** showed the significant alteration in the intensity of XRD peaks intensity after biofield treatment as compared to the control sample. The most intense peak in control sample was observed at 19.06°; while in treated sample the most intense peak was observed at 29.38°. The literature suggests that alteration in crystal morphology may lead to alteration in relative intensities of the

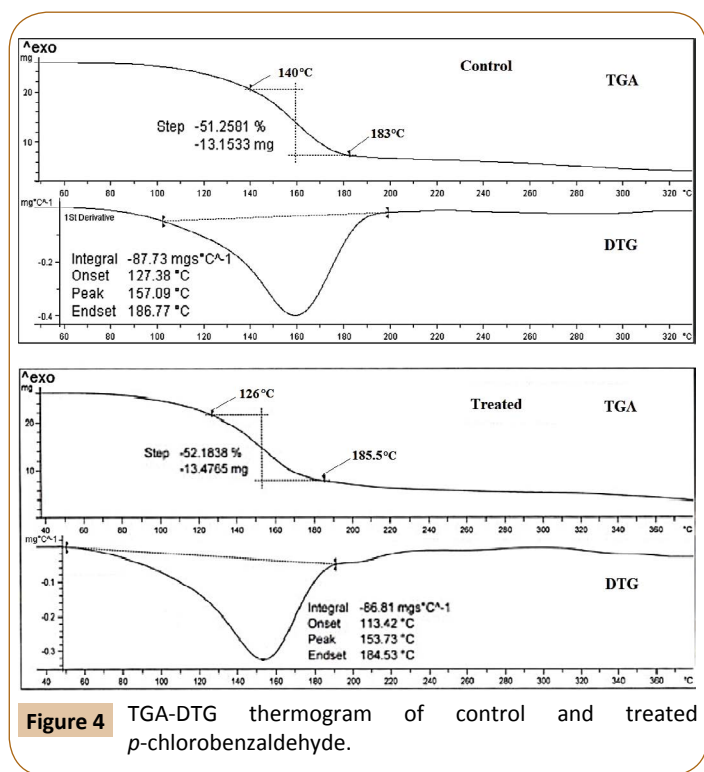
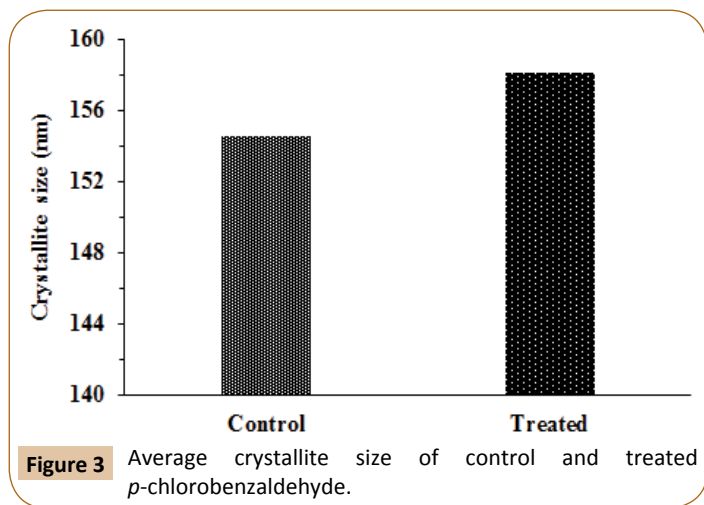
peaks [33]. Additionally, it is reported that internal strain can also change the  $2\theta$  values [34]. Based on this, it is hypothesized that biofield energy treatment was induced an internal strain in the treated sample that might be responsible for the alteration in its  $2\theta$  values with respect to the control sample.

The average crystallite size of the control sample was calculated as 154.52 nm, while the crystallite size of treated sample was calculated as 158.09 nm. The result depicted a slight increase (2.31%) in the crystallite size of treated sample with respect to the control (**Figure 3**). It is previously reported that increase in annealing temperature expressively affects the crystallite size of the compounds. The increase in temperature might lead to decrease in dislocation density and increase in the number of unit cell; these finally increases the average crystallite size of sample [35,36]. Based on this, it is assumed that biofield treatment might provide some thermal energy to *p*-CBA molecules. Consequently, the dislocation density might be reduced and thus the number of unit cells and average crystallite size were increased in the treated sample.

### TGA-DTG analysis

The TGA-DTG thermogram of *p*-CBA samples (control and treated) are shown in **Figure 4** and data are presented in (**Table 1**). The TGA thermogram of control sample showed an initiation (on-set) of thermal degradation at 140°C, which was ended (end-set) at





**Table 1** Thermal analysis of control and treated samples of *p*-chlorobenzaldehyde.  $T_{max}$ : Temperature at maximum weight loss occurs.

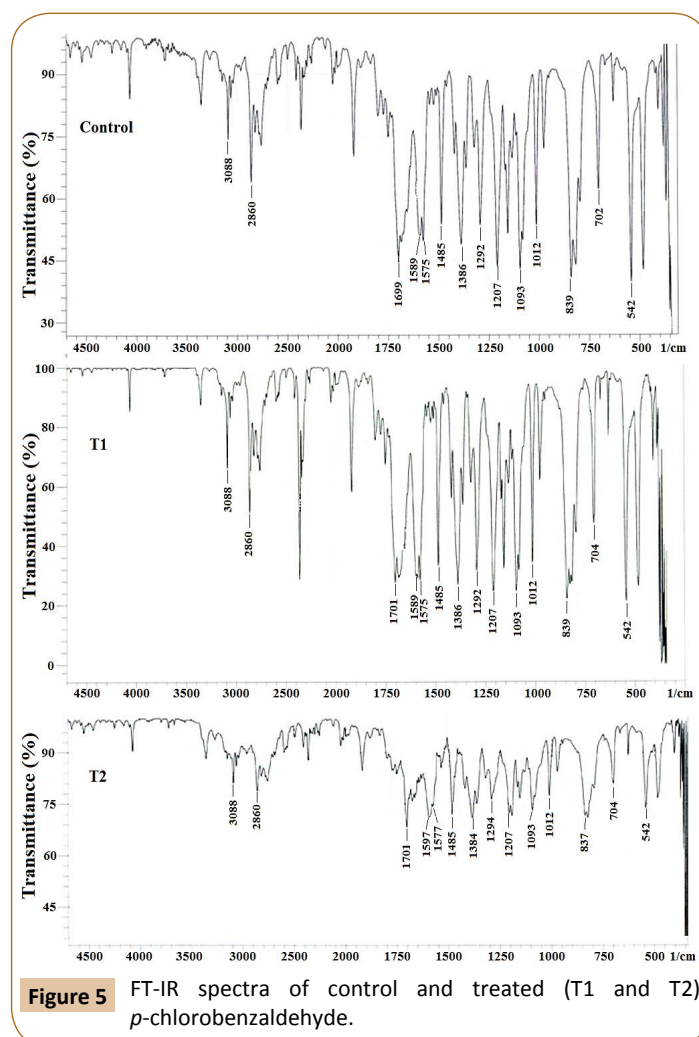
Parameter	Control	Treated
Onset temperature (°C)	140.00	126.00
End-set temperature (°C)	183.00	185.50
$T_{max}$ (°C)	157.09	153.73

183°C. Similarly, the TGA thermogram of the treated sample exhibited the on-set temperature at 126°C that was terminated (end-set) at 185.5°C. The result showed about 10% decrease in the onset temperature in biofield energy treated sample as compared to the control. The TGA-DTG study showed the decrease in thermal stability of treated sample with respect to the control that could be correlated to increase in the chemical reactivity of *p*-CBA. Moreover, the percentage weight loss during

the thermal decomposition was found as 51.25% in the control, while 52.18% in the treated sample. The result showed a mere increase in percent weight loss during thermal decomposition of treated sample with respect to the control. Moreover, the DTG thermogram exhibited the  $T_{max}$  (temperature at which the sample lost its maximum weight) at 157.09°C in the control sample and at 153.73°C in the treated sample. The result showed about 2.14% decrease in  $T_{max}$  of treated sample as compared to the control. This might occur due to the changes in internal energy *via* biofield energy treatment, which may cause to early phase of evaporation in treated sample with respect to the control [37].

### FT-IR spectroscopic analysis

FT-IR spectra of the control and treated *p*-CBA are shown in **Figure 5**. The *p*-CBA molecule contains =C-H, C=C, C-C, C=O, C-Cl groups of vibrations. The =C-H (aromatic) stretching was attributed to peaks at 3088 cm<sup>-1</sup> in control and treated (T1 and T2) samples. While, the aldehyde C-H stretching was assigned to peak appeared at 2860 cm<sup>-1</sup> in all the control and treated samples. The aldehyde C-H asymmetrical bending was attributed to peak at 1485 cm<sup>-1</sup> in all the three samples (control, T1 and T2). Moreover, the aldehyde C-H symmetrical bending was attributed to peaks observed at 1386 cm<sup>-1</sup> in control and T1 sample and 1384 cm<sup>-1</sup> in T2 sample. The out of plane ring deformation was assigned



to peaks at 1093-1207  $\text{cm}^{-1}$  region of spectra in all three samples; whereas, the in-plane deformation was assigned to peaks in the range of 702-839  $\text{cm}^{-1}$  (control), 704-839  $\text{cm}^{-1}$  (T1), and 704-837  $\text{cm}^{-1}$  (T2) sample.

The C=C (aromatic) stretching was assigned to peaks appeared at 1575-1589  $\text{cm}^{-1}$  in control and T1 samples, while it was appeared at 1577-1597  $\text{cm}^{-1}$  in T2 sample. Similarly, the C-C stretching was assigned to peaks at 1292  $\text{cm}^{-1}$  in control and T1 samples, while it was observed at 1294  $\text{cm}^{-1}$  in T2 sample. The C=O stretching peak was assigned to peaks at 1699  $\text{cm}^{-1}$  in control sample and 1701  $\text{cm}^{-1}$  in the treated samples. In addition, the C-O and C-Cl stretching were appeared at 1012  $\text{cm}^{-1}$  and 542  $\text{cm}^{-1}$ , respectively in all the three samples (control, T1 and T2).

The result showed a slight increase in the frequency of C=C stretching in T2 sample as compared to the control. This is might be due to increased bond strength of C=C group in treated *p*-CBA molecules as compared to the control. The stretching frequency of any bond depends on the dipole moment ( $\mu$ ) and reduced mass ( $m$ ) [38,39]. Therefore, it is presumed that biofield energy treatment might increase the dipole moment of C=C bond as compared to the control sample. Except this, rest of the IR vibration peaks were appeared at the similar frequency region in all three samples.

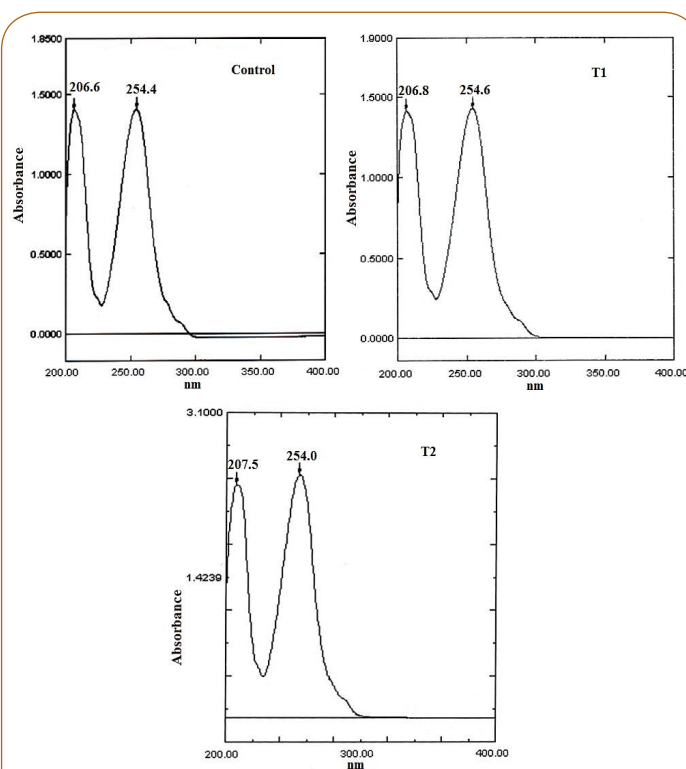
### UV-Vis spectroscopy

UV spectra of the control and treated *p*-CBA are shown in **Figure 6**. The UV spectrum of control sample showed the absorbance maxima ( $\lambda_{\text{max}}$ ) at 206.6 and 254.4 nm. Similarly, the UV spectra of treated sample showed the  $\lambda_{\text{max}}$  at 206.8 and 254.6 nm in T1 and 207.5 and 254.0 nm in T2 sample. The result showed the similar pattern of absorbance maxima in the control and treated samples.

The compound absorbs UV waves due to transition of electrons from highest occupied molecular orbital (HOMO) to highest unoccupied molecular orbital (LUMO). When the energy gap between HOMO and LUMO (also called as HOMO-LUMO gap) altered, the wavelength ( $\lambda_{\text{max}}$ ) was also altered [31]. However, the UV study of *p*-CBA showed the similar pattern of absorbance maxima in both the control and treated samples. Therefore, it can be concluded that the biofield treatment did not distract the energy gap between HOMO-LUMO in treated sample, as compared to the control sample.

### Conclusions

In conclusion, the present study showed the substantial increase in surface area of treated sample by 23.06% as compared to the control sample. The XRD study showed the crystalline nature of both control and treated sample. Moreover, the intensity of XRD peaks were also altered after biofield treatment as compared



**Figure 6** UV spectra of control and treated (T1 and T2) *p*-chlorobenzaldehyde.

to the control. The TGA-DTG study showed the slight decrease (10.0%) in onset temperature of thermal degradation with respect to the control. The decrease in thermal stability might be correlated to increase in chemical reactivity of *p*-CBA. The FT-IR data showed the upstream shifting of C=C stretching frequencies with respect to the control. This might be due to the increase in force constant and bond strength of C=C group in treated *p*-CBA molecule as compared to the control.

Overall, the present study concluded the substantial impact of Mr. Trivedi's biofield energy treatment on physicochemical and spectroscopic properties of *p*-CBA. Based on this, it is anticipated that Mr. Trivedi's unique biofield energy treatment can effectively transform the physicochemical properties of *p*-CBA into the more useful form so that it could be utilized as a better chemical intermediate for the synthesis of pharmaceutical drugs and organic chemicals.

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